

References and Notes

1. C. T. Ewing, J. A. Grand, R. R. Miller, *J. Am. Chem. Soc.* **74**, 11 (1952); *J. Phys. Chem.* **58**, 1086 (1954).
2. C. T. Ewing, J. P. Stone, J. R. Spann, E. W. Steinkuller, D. D. Williams, R. R. Miller, "High Temperature Properties of Sodium and Potassium," *12th Progress Report for U.S. Naval Research Laboratory, NRL Rept. 6094* (Washington, D.C., 9 June 1964).
3. A. W. Lemmon, Jr., H. W. Deem, E. A. Aldridge, E. H. Hall, J. Matolich, Jr., J. F. Walling, "Engineering Properties of Potassium," *NASA CR-54017, BATT-4673-Final* (Battelle Institute, Columbus, Ohio, 1963).
4. I. I. Novikov *et al.*, *J. Nucl. Energy* **4**, 387 (1957).
5. A. V. Grosse, *Science* **140**, 784 (Table 4) (1963).
6. ———, *J. Phys. Chem.* **68**, 3419 (1964).
7. ———, *J. Inorg. Nucl. Chem.* **22**, 23 (1961).
8. E. N. da C. Andrade, *Phil. Mag.* **17**, 698 (1934); both his I and II equations are also fully discussed in ref. 7.
9. M. Sittig, "Sodium, Its Manufacture, Preparation and Uses," *Am. Chem. Soc. Monogr. No. 133* (Reinhold, New York, 1956), pp. 456-461.
10. J. O. Hirschfelder, C. F. Curtiss, R. B. Bird, *Molecular Theory of Gases and Liquids* (Wiley, New York, 1954), p. 14.
11. L. Pauling, *The Nature of the Chemical Bond* (Cornell Univ. Press, Ithaca, N.Y., ed. 3, 1960), p. 403.
12. For mercury, the constant is 83×10^{-4} ; thus it may change to some extent in the various families of the periodic system.
13. O. A. Hougen and K. M. Watson, *Chemical Process Principles* (Wiley, New York, 1943), vol. 3, p. 873.
14. A. V. Grosse, *Inorg. Chem.* **1**, 436 (1962).
15. This work was supported by AEC contract AT(30-1)-2082.

18 January 1965